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DBUH+I₃ complex an efficient catalyst for the synthesis of 2-phenyl benzimidazole and benzothiazole derivatives

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Abstract: Herein, we have reported the facile synthesis of various benzimidazole / benzothiazole by using DBU-Iodine-Iodide as a green and simple catalyst. The R_2NH+I_3 complexes have been formed by reacting an aqueous mixture of ammonium iodide and molecular iodine with the aqueous solution of amine. The structure of R_2NH+I_3 complexes has confirmed by spectroscopic techniques. The prepared amine-iodine complexes have screened as a catalyst in the synthesis of benzimidazole / benzothiazoles. Among the screened catalyst DBUH+I_3 complex has been found as an efficient catalyst. The synthesis of benzimidazoles and benzothiazoles has been achieved with the reaction of *o*-phenylene diamine /*o*- amino thiophenol and various substituted aryl aldehyde using DBUH+I_3 as a catalyst. The present protocol has offered some advantages over other reported protocols such as the mild reaction condition, commercially available precursors, inexpensive catalyst, short reaction time, the broad scope of the substrate, high yield, simple isolation of the product, and environmentally benign method.

Keywords: Amine-iodine complexes; Benzimidazole; Benzothiazole; Oxidative cyclization; organocatalysis

INTRODUCTION

Benzimidazoles and benzothiazoles are valuable heterocyclic scaffolds due to their many applications in diverse fields such as agrochemicals, veterinary, and pharmaceuticals.¹⁻³ They are potent privileged bicyclic aromatic nuclei in organic and medicinal chemistry. They showed diverse biological activity.⁴⁻⁷ Benzimidazole and benzothiazole has found as the core structural skeleton in a variety of drug molecules specifically pantoprazole, riluzole, clemizole, bendamustine, thiabendazole, telmisartan, benzitramide, omeprazole, Hoechst

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33342, pimobendan, Mibefradil, Dovitinib EGFR-3, sulfathiazole, ritonavir, abafungin, tiazofurin, and benazolin. This class of heterocyclic compound displays valuable properties like photochromic, biochemical luminescence, and solvatochromic properties.⁸⁻⁹ These heterocyclic molecules have significant biological activity and great pharmaceutical potential, to attract more attention of synthetic chemists. These heterocyclic molecules have significant biological activity and great pharmaceutical potential, to attract more attention from synthetic chemists (Figure 1).



Figure 1 : Benzimidazoles ring containg drug molecules

The robust method for synthesis of these molecules involves the treatment of *o*-phenylenediamine¹⁰ and 2-amino thiophenol¹¹ with carbonyl compounds, such as aldehyde using Bronsted or Lewis acid catalyst¹² and carboxylic acids¹³ or their derivative (nitrile, amide, ester, acid chloride)¹⁴ at elevated temperature. Another approach involves metal-catalyzed direct alkylation of these molecules via C-H activation followed by carbon-carbon bond formation.¹⁵ Synthesis of these molecules was achieved by microwave,¹⁶ ultrasonic wave,¹⁷ ionic liquid,¹⁸ ionic liquid gel,¹⁹ nanomaterial,²⁰ DMF,²¹ and under oxidative condition using various oxidative and catalytic reagents cited in the reference.²²⁻²³ The certain green synthesis of benzimidazole was accomplished by homogeneous catalysis such as use of triflate erbium catalyst,²⁴ use of active deep eutectic solvent²⁵ and montmorillonite K 10 heterogenous green catalyst.²⁶ Generally, nearly all methods of benzimidazole synthesis have worked for benzothiazole.²⁷ The reported methods have limitations such as harsh reaction conditions, poor yield, high temperature, hazardous and carcinogenic solvent, expensive catalyst, side reaction, slow reaction rate, toxic reagents or tedious workup procedure, and difficulty to isolate the product from the reaction mixture. Consequently, a search for better catalyst, environmentally benign methodology has continued for the economy and



operational simplicity. Our developed amine-iodine complex catalytic procedure is overcoming these problems.

Iodine catalysis has been known for more than 100 years. It has remarkably catalyzed various types of reactions.²⁸⁻²⁹ The drawback of molecular iodine catalyzed synthesis of 2-substituted benzimidazole and benzothiazole is the sublimation of molecular iodine and moisture sensitivity, we have overcome these problems in amine-iodine-iodide complex organocatalyst.

We have synthesized the new R_2NH+I_3 complexes using amine, ammonium iodide, and molecular iodine.³⁰ The R_2NH+I_3 complexes were characterized spectroscopic technique and confirmed.³¹ These catalysts were air-stable, and iodine never sublimates or deliquescent. Amine-iodine complex has catalyzes the synthesis of 2-aryl benzimidazole and benzothiazole, offers several advantages namely short reaction time, easy workup procedure, and environmentally benign protocol. Amine-iodine complexes are organocatalysts that have an indispensable part of synthetic green chemistry because of their stability, less expensive, less toxic, and easily applicable to a wide range of substrates. Herein, we reported amine-iodine complexes catalyzed condensation and cyclization of a wide variety of aryl aldehyde with *o*-phenylenediamine and *o*-amino thiophenol, respectively. Here, we described the synthesis of new amine-iodine complexes (1a-e) and their synthetic application.

EXPERIMENTAL

The Commercially available chemical reagents and solvents were used and their purity was ensured before use. Solvents that were entirely dry and free of impurities were used. Reaction of the progress was checked on Merck TLC Silica gel 60 F254 plates using UV lamp (365 nm and 254 nm) and iodine chamber. The melting point was determined using open capillary method. The recorded melting points were uncorrected. PerkinElmer FTIR spectrometer was used to record IR spectra. Bruker Avance III HD NMR 500 MHz spectrometer is used to obtained ¹HNMR and ¹³C NMR spectra in DMSO d6 and CDCl₃. HRMS analysis was obtained on a Bruker Impact II UHR-TOF mass spectrometer system.

Preparation of DBU-Iodine complexes

2.665 g of Ammonium iodide (18.352 mmol, 2.8 eq.) has added to 5.2 mL water (2 volumes) to got a clear solution in a 250 mL beaker, followed by the addition of 1.667 g of iodine (6.568 mmol, 1 eq). This solution was added dropwise to a stirred solution of 1g DBU (6.568 mmol, 1 eq) in 8 mL water (8 Volume) in a 250 mL round bottom flask. The solid product has formed during addition, stirred the mixture for 15 minutes, and filtered off the solid product. The product has been washed with cold water and dried under a vacuum to provide the desired complexes. After drying the complex, the yield has been reported.

Typical Process for the synthesis of benzimidazole / benzothiazole from o-phenylenediamine/thiophenol and aldehyde.

A mixture of *o*-phenylenediamine/*o*-amino thiophenol (1 mmol) and aryl aldehydes (1 mmol) has dissolved in 2 mL ethanol in a 25 mL round bottom flask. The catalyst DBUH+I₃ complex (1a) (15 mol%) had added to the reaction mixture, and the reaction mixture was stirred



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for 30 min. The progress of the reaction was monitored by (hexane: ethyl acetate) TLC. After completion of the reaction, the solvent has evaporated under a vacuum. The crude reaction mixture was quenched with 20 % sodium thiosulfate solution. The product was isolated by extracting with ethyl acetate. The organic layer was dried over sodium sulfate and purified by column chromatography. The structure of the compound had confirmed by the spectroscopic techniques and matched with the reported.

RESULTS AND DISCUSSION

We have prepared a series of $R_2NH_2 + I_3$ complexes (1a-1e) with minor modification in the reported procedure²⁷⁻²⁸ by replacing potassium iodide with ammonium iodide. This change has led to a drastic change in the structure and composition of catalysts. In the previous reported procedure by Livia et.al.⁷² has formed a precipitate of the complex with composition R_2NH : I_2 : KI. In the present work, we have got a composition as $R_2NH_2+ I_3$ (Scheme 1). Amine must contain two heteroatoms in the cyclic system for precipitation and stability of the complex. The amine like pyrrolidine, piperidine, and amino acid viz proline did not form solid complexes by the same procedure as a result of a single nitrogen atom in the cyclic structure.



Scheme 1 Synthesis of Amine-H-I3 complex and structure of respective complex

The various amine-iodine-iodide complexes have prepared using easily available amine, ammonium iodide, and molecular iodine. The molecular iodine was dissolved in the aqueous solution of ammonium iodide then added to an aqueous solution of amine dropwise, amine-iodine-iodide complex precipitate of respective amine obtained (**Table 1**). The product was washed with excess water till filtrate free from ammonia confirmed by moist turmeric paper.

Table 1 : Synthesis of R₂NH₂+ I₃ complexes^a

No.	Complex	Color	% Yield ^b
1	DBUH+I ₃ complex	Greenish Yellow	92
2	MorpholineH+I3 complex	Orange Yellow	62
3	UrotropineH+I ₃ complex	Brown Yellow	58
4	PiperazineH+I ₃ complexes	Dark Brown Yellow	73
5	N-methyl piperazineH+I3 complexes	Pinkish Yellow	66s

^a:Amine (6.568, 1 equivalent), Iodine (6.568 mmol, 1 equivalent) and Ammonium iodide (18.352 mmol, 2.5 equivalent) in 2mL water ^b: Isolated yield after purification

The structure of synthesized amine-iodine complexes (**1a-1e**) has confirmed by spectroscopic techniques such as UV, IR, HRMS, EDS, ¹HNMR, and ¹³CNMR. These new homogenous catalysts have screened for the synthesis of 2-aryl benzimidazole. We have chosen ethanol as a solvent for screening catalytic activity of the amine-iodine complex catalyst because freely soluble in ethanol and partly soluble in various other organic solvents.

Initially, our studies have being with the screening of prepared amine iodine complexes (1a-1e) for synthesis of benzimidazole, via condensation and cyclization reaction of commercially available *o*-phenylenediamines with *p*-choro benzaldehyde (Scheme 2). The DBUH-I₃ complex has given high yield of 2-(4-chlorophenyl)-*1H*-benzimidazole and the results are given in Table 2.



Scheme 2 Model reaction for screening of R₂NH₂+ I₃ complex for synthesis of benzimidazole

Table 2 : Screening of R_2NH_2 + I_3 complex catalyst in the synthesis of 2-(4-chlorophenyl)-1H-
benzimidazole(4a)^a

Sr. No.	Complex	% Yield ^b
1	DBUH+I ₃ complex	91
2	MorpholineH+I ₃ complex	74
3	UrotropineH+I ₃ complex	85
4	PiperazineH+I ₃ complexes	80
5	N-methyl piperazineH+I ₃ complexes	78
6	Iodine	70
7	Without catalyst	Trace

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^aReaction condition: *o*-phenylenediamine (1 mmol), *p*-chlorobenzaldehyde (1 mmol), $R_2NH_2+I_3$ complex (**1a-1e**) (20 mol %) in ethanol (2 mL) at room temperature for 30 minutes; ^b isolated yield after purification

Next, we have decided to optimize the amount of DBUH+I₃ complex with the same reaction condition. The amount of DBUH-I₃ was optimized by increasing the amount from 5 mol % to 20 mol % for 1 mmol scale reaction. When the reaction has performed in the absence of the catalyst, the product has formed in a very trace amount (**Table 3, entry 1**). The yield has increased with the mol % of amine-iodine complex (**Table 3, entry 2-5**). Nevertheless, there was no increase in the yield when the amount of R_2NH_2 + I₃ catalyst loading has increased from 15 % mol to 20 % mol. From **Table 3**, it has observed, the 15 mol% of DBUH-I₃ complex was sufficient to achieve excellent yield.

Table 3: Optimizing the amount of DBUH+I₃ complex in synthesis of 2-(4-chlorophenyl)-1Hbenzimidazole $(4a)^a$

Entry	Catalyst quantity in mol %	% Yield ^b
1	Without catalyst	Trace
2	5	65
3	10	80
4	15	91
5	20	91

^aReaction condition: *o*-phenylenediamine (1 mmol), *p*-chlorobenzaldehyde (1 mmol), DBUH+ I_3 complex (1a) (mol %) in ethanol (2 mL) at room temperature for 30 minutes; ^b isolated yield after purification.

We have studied the effect of various solvents on product yield (**Table: 4** entry 1-9). Among the screened solvent, ethanol, toluene, and chloroform have given excellent yield, and ethanol has found the best solvent for the reaction as a high amount of product has obtained. Second, fortunately the choice of ethanol also falls on the fact that it is less toxic and more eco-sustainable solvent than chloroform and toluene. Hence, we have selected the solvent for the synthesis of benzimidazole. The solvent DMF, DMSO, and acetonitrile offered a moderate product yield.

 $\label{eq:Table 4: Effect of solvent in synthesis of 2-(4-chlorophenyl)-1H-benzimidazole (4a) using DBUH-I_3 complex catalyst^a$

Entry	Name of solvent	% Yield ^b
1	Ethanol	91
2	Toluene	86
3	Dimethyl formamide	58
4	Dimethyl sulphoxide	66
5	Chloroform	80

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6	Acetic acid	50
7	Acetonitrile	61
8	Tetrahydrofuran	31
9	Water	25

^aReaction condition: *o*-phenylenediamine (1 mmol), *p*-chlorobenzaldehyde (1 mmol), DBUH+I₃ complex (**1a**) (mol %) in ethanol (2mL) at room temperature for 30 minutes; ^b isolated yield after purification

With the investigated optimum reaction condition, we have synthesized various substituted benzimidazole (Scheme 3). The 2-aryl substituted benzimidazole have been synthesized from o-phenylene diamine (1mmol) with several substituted aryl aldehyde (1 mmol) via condensation and cyclization reaction in the presence of DBUH+I₃ complex (15 mol%) at room temperature in ethanol (Table 5). It was found that various substituted aryl aldehyde containing electron-donating groups (*p*-halogen and methoxy, (Table 5 entry 1, 4, 5, 16) and electron-withdrawing group (nitro, Table 5 entry 2, 6, 14) were formed the product with good yield, under optimized condition. The heterocyclic aromatic aldehyde (Table 5, entry 10a, 13a) gave a comparatively lower yield under the same condition. Hydroxy benzaldehyde (Table 5, entry 11, 12) has afforded an unexpectedly low yield, which may be due to solubility in water. The aryl aldehyde bearing electron-withdrawing at ortho/para nitro group (Table 5, entry 13, 15) has afforded product in poor yield. The *o*-substituted aryl aldehyde (Table 5, entry 3, 12, 15) has afforded a low yield due to steric hindrance in cyclization.

Thus, the $R_2NH_2+I_3$ complex was catalyzing the synthesis of 2-aryl substituted benzimidazole using a diverse range of aryl aldehydes and *o*-phenylenediamine. All synthesized benzimidazole derivatives were characterized by ¹H NMR, ¹³C NMR and compared physical constant with standard data. The ¹H NMR displays a characteristic nitrogen-bearing proton chemical shift value 12.5-13.5 δ reflected in each derivative whereas, the ¹³C NMR show a typical chemical shift value 150 δ for carbon located between two nitrogens.



Scheme 3 DBU-Iodine-Iodide catalyzed synthesis of substituted benzimidazole

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Table 5 : Synt	hesis of 2-aryl substituted benzimi	dazole ^a	MD	
Entry	Product (4)	% Yield ^b	М. Р. °С	Literature M. P. °C
1		91°	290- 293	290-292 ²⁹
2		86°	228- 230	227-229 ²⁰
3		73°	232- 234	231-233 ²⁹
4		78 ^d	286- 290	292-293 ²⁹
5		76 ^d	223- 225	222-223 ³⁰
6		81°	202- 205	200-202 ²⁹
7		64 ^d	225- 227	223-226 ³¹
8		72 ^d	238- 240	239-241 ³⁰
9		80°	243- 245	242-244 ²⁰

 Table 5 : Synthesis of 2-aryl substituted benzimidazole^a

	Γ	BUH+I3 CATALYST FOR S	SYNTHESIS		9	
 10		-Me	65 ^d	216- 219	214-216 ²⁰	
11		он	80°	252- 254	254-255 ²⁹	
12		HO 	44°	204- 206	205-206 ³¹	
13			41 ^d	301- 303	30018	
14			72°	196- 198	199 ¹⁸	
15			38°	229- 231	23018	
16			72 ^d	280- 283	277-279 ²⁹	
17		$\square \bigcirc$	51 ^f	270- 273	164-166 ²⁰	
18		\sim	68 ^d	226- 228	221-223 ²⁰	



^aReaction condition: *o*-phenylenediamine (1 mmol), substituted arylaldehyde (1 mmol), DBUH+I₃ complex (**1a**) (15 mol%), EtOH 2 ml, 30 min. at rt; ^bIsolated yield after purification; ^c product was purified by recrystallization in ethanol; ^dproduct was purified by column chromatography mobile phase hexane: ethyl acetate; ^eproduct was purified by recrystallization in chloroform; ^fproduct was purified by column chromatography mobile phase chloroform.





The synthesis of 2-substituted aryl benzothiazole derivative (Scheme 4) has achieved from 2-amino thiophenol and diversity of aryl aldehydes in the presence of DBUH+I₃ complex (1a). The aromatic aldehyde bearing electron-donating group [*p*-halogen, methoxy, hydroxyl, amino, (Table 6 entry 1, 4, 5, 8, 11, 12, 13, and 17)] and electron-withdrawing group (*m*-halogen, methoxy, nitro group,

Table 6 entry 2, 6, 14, 15, 16) provided a good yield of the product under same optimized process. Also, this reaction works well with the heterocyclic aromatic aldehyde to form a product (7) in moderate yield (Table 6, entry 19, 20, 21, 22). The *o*-substituted benzaldehyde has afforded a poor yield of the product because of a steric hindrance (Table 6 entry 3). The unexpectedly *o*-nitro benzaldehyde has afforded a product in the higher yield owing to the high polarity of aldehyde (Table 6 entry 16). Overall, the amine-iodine complex has remarkably catalyzed the synthesis of 2-substituted aryl benzothiazole derivatives. The structure of all synthesized compounds has confirmed by NMR spectroscopic data and compared physical constant with standard data. The ¹³C NMR spectra of benzothiazole have shown a characteristic value of chemical shift 168 δ for carbon between two heteroatoms sulfur and nitrogen.

Ε	ntry	Product (7)	% Yield ^b	M. P. °C	Literature M. P.ºC
	1		84 ^d	115- 117	111-112 ³¹
	2		72°	94-95	93-94 ³¹
	3		58 ^d	80-82	83-84 ³¹
	4	7c N N Td	80°	127- 129	129-131 ³³
	5		74°	120- 121	120-122 ³⁴
	6	CCCSS→OMe 7f	64 ^d	99-102	98-100 ³⁵
	7	CCCSS N 7g ^{OMe} −OMe	61 ^f	229- 231	230-232 ³⁶

Table 6 : Synthesis of 2-aryl substituted benzothiazole^a

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8	S N 7h	83 ^d	130- 132	128-130 ³⁷
9		91 ^d	112- 113	109-110 ³³
10	Ĩ N 7j	62 ^e	85-86	87-88 ³⁸
11	ССС, S N 7k	79°	227- 229	225-227 ³⁹
12		86°	131- 132	124-126 ³⁹
13	ОМе С Л Л Л Л Л Л Л Л Л С М Р ОМе	83°	160- 162	161-163 ³⁹
14		82 ^e	320- 322	228-230 ³⁹
15		78°	190- 193	185-187 ³⁶
16	\sim	80 ^e	195- 197	191-193 ⁴⁰
17		92°	161- 163	160-162 ³⁹



^aReaction condition: o-amino thiophenol (1 mmol), substituted arylaldehyde (1 mmol), DBUH+I₃ complex (**1a**) (15 mol%), Ethanol 2 mL, 30 min. at rt; ^bIsolated yield after purification; ^c product was purified by recrystallization in ethanol; ^dproduct was purified by column chromatography mobile phase hexane: ethyl acetate; ^cproduct was purified by recrystallization in chloroform; ^fproduct was purified by column chromatography mobile phase chloroform.

Further the scope of reaction has extended with the aliphatic aldehydes like crotonaldehyde, propionaldehyde, acetaldehyde, and formaldehyde with *o*-phenylenediamine and *o*-amino thiophenol. The reaction has not proceeded with aliphatic aldehydes and has not afforded the desired product.

Although the exact mechanism is not clear, a proposed mechanism for the formation of benzimidazole and benzothiazole is shown in **Scheme 5**. In first step the aldehyde (**A**) oxygen was protonted by abstraction of proton from DBUH + I_3 complex and form compound (**C**) and liberates DBU + I_3 complex. Simultaneously liberated DBU + I_3 complex, I^- abstract the hydrogen from amines (**B**) to form compound (**D**) and liberates DBU + I_2 complex. In next step (**C**) and (**D**) reacted to form intermediate (**E**). The intermediate (**E**) on reaction DBU + I_2 complex, DBU abstract the proton of XH to form X⁻ and Iodine coordinate with I_2 undergo

cyclization to form intermediate G which undergo oxidative elimination to form C-N double bond to formed final product (**H**).



Scheme 5 Tentative mechanism of DBUH-I₃ catalyzed synthesis of benzimidazole and benzothiazole

CONCLUSION

In the present work, we have prepared the new $R_2NH_2+I_3$ complexes and studied their catalytic activity in the preparation of 2-aryl substituted benzimidazole and benzothiazole derivatives. Among the screened Amine-Iodine catalysis, DBUH+I₃ has found an efficient catalyst for the preparation of 2-aryl substituted benzimidazole and benzothiazole. We have believed that the present method is more convenient, efficient, greener, simple, and environmentally benign than most reported methods in the Literature. The present method has not afforded the benzimidazole and benzothiazole derivatives with aliphatic aldehydes.

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SUPPLEMENTARY MATERIAL

Additional data are available electronically at the pages of journal website: <u>https://www.shd-pub.org.rs/index.php/JSCS/article/view/11893</u>, or from the corresponding author on request.

ИЗВОД

КОМПЛЕКС DBUH+I3 КАО ЕФИКАСАН КАТАЛИЗАТОР ЗА СИНТЕЗУ ДЕРИВАТА 2-ФЕНИЛБЕНЗИМИДАЗОЛА И БЕНЗОТИАЗОЛА

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У овом раду је описана једноставна синтеза различитих бензимидазола/бензотиазола, употребом DBU-јод-јодида као једноставног и еколошки прихватљивог катализатора. Настаје комплекс R₂NH+I₃ у реакцији смеше амонијум јодида, молекулског јода и амонијака у води. Структура комплекса R₂NH+I₃ потврђена је спектроскопским техникама. Каталитичке особине добијеног амин-јодидног комплекса су испитане у реакцији синтезе бензимидазола/бензотиазола. Од испитаних катализатора DBUH+I₃ комплекс се показао као ефикасан. Синтеза бензимидазола и бензотиазола је постигнута у реакцијама *о*фенилендиамина /*о*- аминотиофенола са различитим супституисаним арил-алдехидима користећи DBUH+I₃ комплекс као катализатор. У односу на друге, приказани протокол има неколико предности, као што су благи реакциони услови, комерцијално доступни прекурсори, катализатор који није скуп, кратко реакционо време, ширк опсег супстрата, висок принос, једноставан поступак изоловања производа, и поступак који није штетан за животну средину.

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SUPPLEMENTARY MATERIAL TO DBUH+I₃ complex an efficient catalyst for the synthesis of 2-phenyl benzimidazole and benzothiazole derivatives

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General:

All local brand chemicals were purchase checked their purity by TLC and purified. The melting point were determined in open capillary and are uncorrected. For analysis technique following instruments were used. Solvents that were entirely dry and free of impurities were used. Reaction of the progress was checked on Merck TLC Silica gel 60 F254 plates using UV lamp (365 nm and 254 nm) and iodine chamber.

Sr. No.	Analysis Type	Instrument
1	HRMS	Brucker Impact HD
2	UV-visible Spectrum	shimadzuCorp, Model UV-2600
3	IR Spectrum	shimadzu Corp, FTIR-shimay, Model IR affinity
4	FESM	FEI Nova NanoSEM 450
5	EDS	Brucker XFlash 6130
6	TGA-DTA	shimadzu Corp
7	NMR (1 H & 13 C)	500MHz &125MHZ Brucker

Synthesis of amine-iodine complexes

Ammonium iodide (2.8 eq.) was added to water (2 volume) has obtained clear solution in 250 mL beaker and then added iodine (1 eq). This mixture of the solution was added dropwise to a stirred solution of amine (1 eq) in water (8 Volume) in 250 mL round bottom flask. The solid product has formed during addition, stirred mixture for 15 minutes and filter off the solid product. The product was washed with cold water and dried under vacuum to provide the desired complexes and the yield of the complex was reported.

Typical Process for the synthesis of benzimidazole / benzothiazole from ophenylenediamine/thiophenol and aldehyde.

A mixture of *o*-phenylenediamine/*o*-amino thiophenol (1 mmol) and arylaldehyde (1 mmol) was dissolved in 2 mL ethanol in 25 mL round bottom flask. The catalyst (**1a**) (15mol%) was added and the reaction mixture was stirred for 30 min. The progress of the reaction was monitored by (hexane: ethyl acetate) TLC. The TLC clearly have showed the disappearance of thestarting material. After completion of the reaction, the solvent was evaporated under vacuum. The crude solid product was extracted in ethyl acetate after the addition of 20 % sodium thiosulphate solution. The organic layer was dried over sodium sulfate and purified by column chromatography. The structure of the compound was confirmed by the spectroscopic techniques and match with the reported.

1a. DBU-Iodine complex (Table 1, Entry 1, 1a): Greenish Yellow solid M. P. 87^oC.



SUPPLEMENTARY MATERIAL



M. F. = $C_9H_{17}N_2 I^- I_2$ Mol. Wt. = 533.79

HRMS: Positive ion polarity: 153.138 (cal. 153.242).

Negative ion polarity: 126.904 (cal. 126.904), 380.712 (cal. 380.713).

UV-visible Spectrum(nm): 210, 307,364 ($\lambda_{max} = 364$ nm).

IR Spectrum(cm⁻¹): 530, 601, 633, 1203, 1319, 1440, 1574, 1638, 3133, 3267. **SEM**:Clumpy and agrummerated morphology.

Field Emission Scanning Electron Microscopy Energy Dispersive X-ray **Spectroscopy (FESEM - EDS):**

Element	At. Number	Wt. %	At. %
Iodine	53	78.97	26.69
Carbon	6	17.57	62.74
Nitrogen	7	3.45	10.57
		100	100

TGA: DBU-iodine complex was stable up to 200°C after that gradual weight loss start up to 380°C then fast weight loss observed and stop 410°C. After 410°Cslow weight loss starts and end by complete vanishing of complex at 500°C.

DTA: Endotherm was observed at 110°C and exotherm at 410°C. Both peaks are very sharp.

Strong exotherm and sharp weight loss was located in graph at 410°C.

¹HNMR:(500 MHz,DMSO-d₆):δ9.47 (s, 1H), 3.55 (t, 2H J=3.55 Hz), 3.48 (t, 2H J=3.48 Hz), 3.24-3.26 (m, 2H), 2.63 (t, 2H J=2.64 Hz), 1.92(q, J=1.94Hz) 1.54-1.72 (m, 6H);¹³CNMR:(125 MHz,DMSO-d₆) δ:165.88, 53.89, 48.38, 38.10, 32.22, 28.70, 26.38, 23.78, 19.34.

1b. Morpholine-Iodine complex (Table 1, Entry 2, 1b): Orange Yellow solid M. P. 78°C.



M. F. = $C_4H_9NO I^-I_2$ Mol. Wt. = 467.73

HRMS: Positive ion polarity: 88.075 (cal. 88.126). Negative ion polarity: 126.905 (cal. 126.904), 380.713 (cal. 380.713). UV-visible Spectrum(nm): 210, 360, 365, 366. ($\lambda_{max} = 360$ nm).

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IR Spectrum (cm⁻¹): 585, 626, 817, 859, 1006, 1033, 1083, 1159, 1243, 1295, 1357, 1438, 2858, 3183.

SEM: Clumpy and agrummerated morphology.

Field Emission Scanning Electron Microscopy Energy Dispersive X-ray Spectroscopy (FESEM - EDS):

Element	At. Number	Wt. %	At. %
Iodine	53	66.21	16.45
Carbon	6	23.94	62.86
Oxygen	8	5.36	10.57
Nitrogen	7	4.49	10.11
		100	100

TGA:The morpholine-iodine complex was stable up to 150°C then underwent fast weight loss till 300°Cthen gradual weight loss observed end at 500°Cby complete disappearing complex.

DTA: It displays harp endotherm at 150°C and very broad exotherm peak at 480°C.

¹HNMR: (500 MHz, DMSO-d₆):δ 3.66-3.77 (m, 4H), 3.35-3.41(m,

4H);¹³CNMR:(125 MHz, DMSO-d₆)δ:44.09, 45.24, 64.46, 65.55.

1c. Urotropine-Iodine complex (Table 1, Entry 3, 1c): Brown Yellow solid M. P. 130^oC.



M. $\mathbf{F}_{.} = \mathbf{C}_{6}\mathbf{H}_{13}\mathbf{N}_{4}\mathbf{I}^{-}\mathbf{I}_{2}$ **Mol.** Wt. = 521.76

HRMS: Positive ion polarity: 141.113 (cal. 141.192).

Negative ion polarity: 126.905 (cal. 126.904).

UV-visible Spectrum(nm): 308, 113, 324, 369 ($\lambda_{max} = 369$ nm).

IR Spectrum (cm⁻¹): 523, 656, 705, 734, 819, 901, 991, 1028, 1230, 1250, 1381, 1455.

SEM:Clumpy and agrummerated morphology.

Field Emission Scanning Electron Microscopy Energy Dispersive X-ray Spectroscopy (FESEM - EDS):

Element	At. Number	Wt. %	At. %
Iodine	53	74.50	22.60
Carbon	6	16.03	51.37
Nitrogen	7	09.47	26.03
		100	100

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TGA: The complex was very stable up to 255°C after that sharp decrease in weight continue till temperature 340°C.

DTA: It show sharp three band at temperature 145°C, 255°Cand 450°C. ¹HNMR:(500 MHz, DMSO-d₆):64.73(s, 12H);¹³CNMR:(125 MHz, DMSOd₆)δ:73.85.

1d. Piperazine-Iodine complex (Table 1, Entry 4, 1d): Dark Brown Yellow Solid M. P. 346^oC.



M. F. = $C_4H_{11}N_2 I^- I_2$ Mol. Wt. = 467.74

HRMS: Positive ion polarity: 87.091(cal. 87.142).

Negative ion polarity: 126.905 (cal. 126.904).

UV-visible Spectrum(nm): 210, 306, 319, 361, 368 ($\lambda_{max} = 368$ nm).

IR Spectrum (cm⁻¹): 636, 860, 988, 1084, 1242, 1358, 1400, 1436, 3180. SEM:Clumpy and agrummerated morphology.

Field Emission Scanning Electron Microscopy Energy Dispersive X-ray **Spectroscopy (FESEM - EDS):**

 Element	At. Number	Wt. %	At. %
Iodine	53	89.50	45.67
Carbon	6	07.55	40.73
Nitrogen	7	02.94	13.60
		100	100

TGA: The complex show stability till temperature 115°Cafter sharp and slow weight loss continue up to 325°C.

DTA: This graph indicates one sharp exothermic band at 325°C.

¹HNMR: (500 MHz, DMSO-d₆): δ8.48(s, 2H), 3.81 (s, 1H), 3.22(t, 1H, J=3.22 Hz), 3.07 (t, 4H, J= 3.08Hz), 2.99 (s, 1H), 2.61-2.64 δ (q, 1H);¹³CNMR: (125) MHz, DMSO-d₆) δ:47.85, 46.74, 44.28, 43.54.

1e. N-Methyl-Piperazine-Iodine complex (Table 1, Entry 5, 1e): Pinkish Yellow solid M. P. 178°C.



M. F. = $C_5H_{13}N_2 I^- I_2$ Mol. Wt. = 481.75

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HRMS: Positive ion polarity: 101.107 (cal. 101.168). Negative ion polarity: 126.905 (cal. 126.904). UV-visible Spectrum(nm): 210, 306, 317, 365 ($\lambda_{max} = 364$ nm).

IR Spectrum (cm⁻¹): 573, 847, 893, 960, 990, 1100, 1365, 1438, 1553, 1651, 2436, 2707.

SEM:Clumpy and agrummerated morphology.

Field Emission Scanning Electron Microscopy Energy Dispersive X-ray Spectroscopy (FESEM - EDS):

Element	At. Number	Wt. %	At. %
Iodine	53	76.30	23.96
Carbon	6	18.18	60.32
Nitrogen	7	05.52	15.71
		100	100

TGA: The complex was stable till 200°C above this temperature gradual weight loss till 320°C.

DTA: One sharp exothermic band observed at 320°C.

¹HNMR:(**500** MHz, DMSO-d₆): δ8.46 (s, 2H), 2.96-3.05(m, 4H), 2.61-2.63(m, 4H), 2.35(s, 3H);¹³CNMR:(**125** MHz, DMSO-d₆) δ:51.47, 45.40, 43.02.

CHARACTERISATION DATA OF 2-SUBSTITUTED PHENYL BENZIMIDAZOLE.

1. 2-(4-chlorophenyl)-*1H***-benzimidazole (Table 5, Entry 1, 4a):** Yellow solid M. P. 290-293°C (290-292°C)¹



¹HNMR: (500 MHz, DMSO-*d*₆): δ 12.98 (s, 1H), 8.17-8.20 (m, 2H), 7.73 (d, 1H, *J*=7.73 Hz), 7.64-7.68 (m, 1H), 7.63 (t, 1H, *J*= 7.62 Hz), 7.61 δ (d, 1H *J*= 7.60 Hz), 7.20-7.36 (m, 2H); ¹³CNMR: (125 MHz, DMSO-*d*₆)δ: 150.61, 144.20, 135.48, 134.95, 129.54, 129.27, 128.60, 123.24, 122.31, 119.43, 111.88. **2. 2-(3-chlorophenyl)-1***H***-benzimidazole (Table 5, Entry 2, 4b):** Brown solid M. P. 228-230°C (227-229°C)²



¹HNMR: (500 MHz, DMSO- d_6): δ 13.04 (s, 1H), 8.23 (t, 1H J= 8.22 Hz), 8.17 (t, 1H J= 8.17 Hz), 8.15 (t, 1H, J= 8.13 Hz) 7.57-7.66 (m, 1H), 7.55 (t, 1H, J= 7.54 Hz), 7.30 (q, 1H), 7.20 – 7.27 δ (m, 2H); ¹³CNMR: (125 MHz, DMSO-

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 d_6) δ : 150.19, 144.11, 135.45, 134.23, 132.67, 131.42, 130.01, 126.48, 125.48, 123.43, 122.41, 119.56, 111.98.

3. 2-(2-chlorophenyl)-*1H*-benzimidazole (Table 5, Entry 3, 4c): Yellow solid M. P. 232-234°C (231-233⁰C)¹



¹HNMR: (500 MHz, DMSO-*d*₆): δ 12.73 δ (s, 1H), 7.90-7.91 (m, 1H), 7.66 (d, 1H, *J*= 7.65 Hz), 7.65 (d, 2H, *J*=7.65 Hz), 7.50-7.56 (m, 2H), 7.22-7.26 (m, 2H); 1³CNMR: (125 MHz, DMSO-*d*₆)δ :149.55, 132.56, 132,09, 131.68, 130.82, 130.43, 127.91, 122.72, 120.07.

4. 2-(4-bromophenyl)-*1H*-benzimidazole (Table 5, Entry 4, 4d): Yellow solid M. P. 286-290°C (292-293°C)¹



5. 2-(4-Methoxyphenyl)-*1H*-benzimidazole (Table 5, Entry 5, 4e): White solid M. P. 223-225°C (222-223°C)³



¹HNMR: (500 MHz, DMSO-*d*₆): δ 12.73 (s, 1H), 8.10 - 8.12 (m, 2H), 7.61 (d, 1H, J= 7.60 Hz), 7.48 (d, 1H, J=7.48 Hz), 7.17 (t, 2H, J= 7.16 Hz) 7.13 (d 1H J=7.13 Hz) 7.11 (d, 1H J=7.10 Hz) 3.84 (s, 3H); ¹³CNMR: (125 MHz, DMSO-*d*₆)δ: 161.05, 151.79, 144.34, 135.43, 128.45, 123.15, 122.53, 121.90, 118.95, 114.83, 111.49.

6. 2-(3-methoxyphenyl)-1*H***-benzimidazole (Table 5, Entry 6, 4f):**Yellow solid M. P. 202-205°C (200-202°C)¹



7. 2-(3,4-dimethoxyphenyl)-*1H*-benzimidazole (Table 5, Entry 7, 4g): White solid M. P. 225-227°C (223-226°C)⁴

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8. 2-(2H-1,3-benzodioxol-5-yl)-*1H*-benzimidazole (Table 5, Entry 8, **4h**):Yellow solid 238-240°C (239-241°C)³



¹HNMR: (500 MHz, DMSO-*d*₆): δ 12.80 (s, 1H), 7.67 (q, 1H, *J*= 7.66 Hz), 7.45-7.47 (m, 1H), 7.27 (d, 1H, *J*= 7.26 Hz), 7.21-7.23 (m, 3H), 6.59(d, 1H, *J*= 6.58 Hz), 5.96 (s, 2H); ¹³CNMR: (125 MHz, DMSO-*d*₆)δ: 153.50, 148.07, 147.04, 143.02, 136.25, 124.22, 123.00, 122.60, 119.85, 119.55, 111.53, 108.91, 107.22, 101.58.

9. 2-phenyl-*1H***-benzimidazole (Table 5, Entry 9, 4i):** Brown solid M. P. 243-245°C (242-244°C)²



¹HNMR: (500 MHz, DMSO-*d*₆): δ 12.91 (s, 1H), 8.19 (t, 2H *J*=8.18 Hz), 7.67 (d, 1H, *J*=7.67 Hz), 7.53-7.57 (m, 3H), 7.50 (t, 1 H *J*=7.48 Hz) 7.18-7.24 (m, 2H); ¹³CNMR: (125 MHz, DMSO-*d*₆)δ: 151.68, 144.28, 135.47, 130.64, 130.30, 129.41, 129.25, 127.09, 126.90, 122.99, 122.13, 119.34, 111.78.

10. 2-(4-methylphenyl)-*1H*-benzimidazole (Table 5, Entry 10, 4j):Brown Solid M. P. 216-219°C (214-216°C)²



11. 4-(*1H***-benzimidazole-2-yl) phenol (Table 5, Entry 11, 4k):** White solid M. P. 252-254°C (254-255°C)¹



¹HNMR: (500 MHz, DMSO- d_{δ}): δ 15.33 (s, 1H), 10.87 (s, 1H), 8.25 (d, 2H J= 8.24 Hz) 7.77-7.81 (m, 2H), 7.51-7.54 (m, 2H), 7.09-7.11 (d, 2H, J=7.09 Hz);

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¹³CNMR: (125 MHz, DMSO-*d*₆)δ: 162.86, 149.60, 132.12, 130.74, 125.97, 116.98, 114.04, 113.78.

12. 2-(*1H***-benzimidazole-2yl) phenol (Table 5, Entry 12, 4l):**Brown solid 204-206°C (205-206°C)⁵



¹HNMR: (500 MHz, DMSO-*d*₆): δ 13.29 (s, 1H), 13.09 (s, 1H), 8.06 (d, 1H, *J*=8.05 Hz), 7.94-7.97 (m, 2H), 7.90 (d, 1H, *J*=7.89Hz), 6.61-7.64 (m, 2H), 7.48-7.51(m, 1H), 7.38-7.41(m, 1H); ¹³CNMR: (125 MHz, DMSO-*d*₆)δ: 156.85, 152.58, 142.33, 131.85, 128.83, 127.16, 123.10, 122.99, 119.50, 116.88, 115.50, 111.29.

13. 2-(4-nitrophenyl)-*1H*-benzimidazole (Table 5, Entry 13 4m): Yellow solid M. P. 301-303°C (300°C)⁶



14. 2-(3-nitrophenyl)-*1H*-benzimidazole (Table 5, Entry 14, 4n): Yellow solid M. P. 196-198°C (199°C)⁶



15. 2-(2-nitrophenyl)-*1H*-benzimidazole (Table 5, Entry 15, 40): Yellow solid M. P. 229-231°C (230°C)⁶



¹HNMR: (500 MHz, DMSO-*d*₆): δ 13.06 (s, 1H), 8.03 (dd, 1H, *J*=8.02 Hz), 7.98 δ (dd, 1H, *J*=7.97 Hz) 7.85-7.88 (m, 1H), 7.74-7.77 (m, 1H) 7.69 (d, 1H, *J*=7.65 Hz), 7.57 (d, 1H *J*=7.56 Hz) 7.20-7.29 (m, 2H); 13CNMR: (125 MHz, DMSO-*d*₆)δ:149.42, 147.76, 144.05, 135.07, 133.12, 131.38, 124.77, 124.67, 123.56, 122.36, 119.71, 112.14.

16. 4-(*1H*-benzimidazole-2-yl)-N, N-dimethylaniline (Table 5, Entry 16,4p):Yellow solid M. P. 280-283°C (277-279°C)¹

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¹**HNMR:** (500 MHz, CDCl₃): 12.51 (s, 1H), 7.61-7.64 (m, 1H), 7.23-7.26 (m, 1H), 7.15-7.20 (m, 1H), 7.01 (d, 1H, *J*=7.01 Hz), 6.67-6.74 (m, 2H), 3.00 (m, 3H), 2.92 (m, 3H); ¹³CNMR: (125 MHz, CDCl₃)δ: 155.04, 149.98, 143.30, 136.38, 130.31, 126.94, 124.34, 122.16, 119.26, 117.39, 112.81, 111.81, 110.40, 40.57, 40.23.

17. 2-[(*E***)-2-phenylethenyl]-***1H***-benzimidazole (Table 5, Entry 17 4q): Yellow solid M. P. 270-273°C (164-166°C)²**



18. 2-(furan-2-yl)-*1H***-benzimidazole (Table 5, Entry 18, 4r):**Brown solid M. P. 226-228°C(221-223°C)²



¹HNMR: (500 MHz, DMSO-*d*₆): δ 12.95(m, 1H), 7.96(dd, *J*=1.71&0.90Hz, 1H), 7.57(d, *J*=7.11Hz, 1H), 7.51 (d, *J*=7.15Hz, 1H) 7.17-7.22 (m, 3H), 6.72(dd, *J*=3.4 Hz &0.95 Hz, 1H); ¹³CNMR: (125 MHz, DMSO-*d*₆) δ:147.12, 143.37, 135.00, 134.44, 129.02, 128.68, 127.15, 123.01, 121.90, 117.81, 112.59. **19. 3-(1***H***-benzimidazole-2-yl)2-chloroquinoline (Table 5, Entry 19, 4s)**:Yellow solid M. P. 219-222°C (202°C)⁷



¹HNMR: (500 MHz, DMSO-*d₆*): δ 10.41 (s, 1H), 9.34 (s, 1H), 8.08 (d, 1H, *J*=8.06 Hz), 8.06 (d, 1H, *J*=8.05Hz), 7.98-8.00 (m, 1H), 7.81-7.85 (m, 1H), 7.35-7.60 (m, 2H); ¹³CNMR: (125 MHz, DMSO-*d₆*) δ: 147.57, 147.02, 145.58, 143.07, 141.48, 133.96, 131.98, 128.41, 128.30, 128.09, 127.05, 124,12, 123.22, 122,74, 119.86, 111.29.

20. 3-(*1H***-benzimidazol-2-yl)-2-chloro-6-methylquinoline (Table 5, Entry 20, 4t):**White solid M. P. 221-224°C (220°C)⁷



¹HNMR: (500 MHz, DMSO-*d*₆): δ 12.94 (s, 1H), 8.87 (s, 1H), 8.59 (s, 1H), 7.75-7.78 (m, 2H), 7.61-7.74 (m, 2H), 7.28-7.59 (m, 2H), 2.63 (s, 3H); ¹³CNMR: (125 MHz, DMSO-*d*₆) δ: 148.31, 146.73, 145.95, 141.20, 138.32, 134.71, 127.93, 127.60, 126.83, 124.87, 123.53, 122.41, 119.66, 112.32, 21.64. 21. 2-(1*H*-indol-2-yl)-*1H*-benzimidazole (Table 5, Entry 21, 4u):Black solid 220-223°C (226-227°C)¹⁷



¹**HNMR**:(500 MHz DMSO-*d₆*): δ 12.59 (s, 1H), 11.66 (s, 1H), 8.49 (t, 1H, *J*= 8.48 Hz), 8.14 (d, 1H, *J*=8.13 Hz), 7.49-7.55 (m, 2H), 7.54-7.55 (m, 2H), 7.49-7.50 (q, 1H), 7.19-7.21 (m, 1H), 7.13-7.16 (m, 1H); ¹³**CNMR**:(125 MHz DMSO*d₆*) δ: 149.84, 136.96, 126.77, 125.54, 122.73, 121.78, 120.78, 112.41, 106.74. **22. 3-(***1H***-benzimidazol-2-yl)-6-bromo-4***H***-1-benzopyran-4-one (Table 5, Entry 22, 4v): Yellow solid M. P. 269-271°C.**



¹HNMR: (500 MHz, DMSO- d_6): δ 12.65 (s, 1H), 8.32 (s, 1H), 9.41 (s, 1H), 8.32 (d, 1H, J= 8.31 Hz), 8.06-8.09 (m, 1H), 7.67-7.70 (m, 1H), 7.62-7.66 (m, 1H), 7.18-7.22 (m, 2H); ¹³CNMR: (125 MHz, DMSO- d_6) δ : 174.11, 158.91, 155.05, 145.27, 142.70, 137.89, 134.93, 127.82, 125.57, 122.70, 122.40, 121.96, 119.22, 118.73, 115.03, 112.97.

HRMS: [MF: C₁₆H₁₀O₂N₂Br(M+H)]: 342.99 (Calculated: 342.16)

CHARACTERISATION DATA OF 2-SUBSTITUTED PHENYL BENZOTHIAZOLE.

1. 2-(4-chlorophenyl)-1,3-benzothiazole (Table 6, Entry 1, 7a):White solid M. P. 115-117°C (111-112°C)⁸



2. 2-(3-chlorophenyl)-1,3-benzothiazole (Table 6, Entry 2. 7b):White solid M. P. 94-95°C (93-94°C)⁸



3. 2-(2-chlorophenyl)-1,3-benzothiazole (Table 6, Entry 3, 7c): White solid M. P. 80-82°C (83-84°C)⁸



4. 2-(4-bromophenyl)-1,3-benzothiazole (Table 6, Entry 4, 7d): White Solid M. P. 127-129°C (129-131°C)⁹



¹HNMR: (500 MHz, CDCl₃): δ 8.06 (d, 1H, *J*=8.05 Hz), 7.94-7.91 (m, 2H), 7.90(d, 1H, *J*= 7.89 Hz), 7.61-7.64(m, 2H), 7.48-7.51 (m, 1H), 7.38-7.41 (m, 1H); ¹³CNMR: (125 MHz, CDCl₃) δ : 166.70, 154.06, 135.03, 132.54, 132.23, 128.90, 126.51, 125.45, 125.42, 123.31, 121.67.

5. 2-(4-methoxyphenyl)-1,3-benzothiazole (Table 6, Entry 5, 7e):White Solid M. P. 120-121°C (120-122°C)⁹



¹HNMR: (500 MHz, CDCl₃): δ 8.02-8.04 (m, 3H), 7.86 (d, 1H, *J*=7.86 Hz), 7.44-7.48 (m, 1H), 7.33-7.36 (m, 1H), 6.98-7.01 (m, 2H), 3.87 (s, 3H); ¹³CNMR: (125 MHz, CDCl₃) δ: 167.86, 161.91, 154.22, 134.85, 129.10, 126.43, 126.19, 124.78, 122.81, 121.50, 114.36, 55.46.

6. 2-(3-methoxyphenyl)-1,3-benzothiazole (Table 6, Entry 6, 7f):Yellow solid M. P. 99-102°C (98-100°C)¹⁰

SUPPLEMENTARY MATERIAL



7. 2-(3,4-dimethoxyphenyl)-1,3-benzothiazole (Table 6, Entry 7, 7g):Brown solid M. P. 229-231°C (230-232°C)¹¹



8. 2-(2*H*-1,3-benzodioxol-5-yl)-1,3-benzothiazole (Table 6, Entry 8, 7h):Yellow solid M. P. 130-132°C (128-130°C)¹²



¹HNMR: (500 MHz, CDCl₃): δ 8.00 (d, 1H, *J*=7.99 Hz), 7.81 (d, 1H, *J*=7.81 Hz), 7.57 (d, 1H, *J*=7.56 Hz), 7.42-7.45 (m, 1H), 7.30-7.35 (m, 1H), 6.85 (d, 1H, *J*=6.84 Hz), 5.99 (s, 2H); ¹³CNMR: (125 MHz, CDCl₃) δ: 167.49, 154.01, 150.01, 148.29, 134.80, 127.94, 126.20, 124.89, 122.86, 122.66, 122.43, 108.56, 107.43, 101.67.

9. 2-phenyl-1,3 benzothiazole (Table 6, Entry 9, 7i): White solid M. P. 112-113°C (109-110°C)⁸



10. 2-(4-methylphenyl)-1, 3-benzothiazole (Table 6, Entry, 10, 7j):Yellow solid M. P. 85-86°C (87-88°C)¹³



11. 4-(1,3-benzothiazol-2-yl) phenol (Table 6, Entry 11, 7k): White solid M. P. 227-229°C (225-227°C)¹⁴



GAWADE and KULKARNI.

¹HNMR: (500 MHz DMSO-*d*₆): δ 10.24 (s, 1H), 8.09 (d, 1H, *J*=8.07), 8.07 (d, 1H, *J*=8.06), 7.93-7.00 (m, 2H), 7.49-7.52 (m, 1H), 7.39-7.42 (m, 1H), 6.96 (t, 2H, *J*=6.95); ¹³CNMR: (125 MHz DMSO-*d*₆) δ : 167.92, 160.99, 154.19, 134.57, 129.74, 129.51, 129.27, 126.89, 125.36, 124.50, 122.76, 122.58, 116.55. **12. 2-(1,3-benzothiazol-2-yl) phenol (Table 6, Entry 12, 7l):** White solid M. P. 131-132°C (124-126°C)¹⁴



¹HNMR: (500 MHz, CDCl₃): δ 12.50 (s, 1H), 7.97 (d, 1H, *J*=7.96 Hz), 7.87 (d, 1H, *J*=7.87 Hz), 7.66-7.68 (m, 1H), 7.47-7.50 (m, 1H), 7.35-7.70 (m, 2H), 7.09-7.10 (m, 1H), 6.92-6.96 (m, 1H); ¹³CNMR: (125 MHz, CDCl₃) δ: 169.35, 157.92, 151.81, 132.73, 132.56, 128.39, 128.14, 126.66, 125.52, 122.16, 121.49, 119.53, 117.85.

13. 4-(1,3-benzothiazol-2-yl)-2-methoxyphenol (Table 6, Entry 13, 7m):White solid M. P. 160-162°C (161-163°C)¹⁴



¹**HNMR:** (500 MHz, CDCl₃): δ 8.03 (d, 1H, *J*= 8.01 Hz), 7.86 (q, 1H), 7.71 (d, 1H, *J*= 7.70 Hz), 7.54 (q, 1H), 7.45-7.48 (m, 1H), 7.33-7.37 (m, 1H), 7.00 (q, 1H), 6.10 (s, 1H) 4.00(s, 3H); ¹³CNMR: (125 MHz, CDCl₃) δ : 168.15, 154.04, 148.52, 146.95, 134.81, 126.22, 126.17, 124.84, 122.72, 121.94, 121.51, 114.71, 109.24, 56.17.

14. 2-(4-nitrophenyl)-1,3-benzothiazole (Table 6, Entry14, 7n):Brown solid M. P. 320-322°C (228-230°C)¹⁴



¹HNMR: (500 MHz, CDCl₃): δ 8.92 (s, 1H), 8.40 (d, 1H, J=8.40 Hz), 8.31 (d, 1H, J=8.30 Hz), 8.11 (d, 1H, J=8.10 Hz), 7.94 (d, 1H, J=7.93 Hz), 7.67 (t, 1H, J=7.68 Hz), 7.56 (t, 1H, J=7.56 Hz), 7.45 (t, 1H, J=7.45 Hz); ¹³CNMR: (125 MHz, CDCl₃) δ: 164.89, 153.93, 148.74, 135.17, 133.01, 130.12, 126.85, 126.05, 125.19, 123.75, 122.69, 122.32, 121.85.

15. 2-(3-nitrophenyl)-1,3-benzothiazole (Table 6, Entry 15, 70):Yellow solid M. P. 190-193°C (185-187°C)¹⁵





16. 2-(2-nitrophenyl)-1,3-benzothiazole (Table 6, Entry 16, 7p):Orange brown solid M. P. 195-197°C (191-193°C)¹⁵



¹HNMR: (500 MHz, CDCl₃): δ 8.08 (d, 1H *J*=8.07 Hz), 7.88-7.94 (m, 2H) 7.79 (q, 1H), 7.67-7.70 (m, 1H), 7.61-7.64 (m, 1H), 7.51-7.54 (m, 1H), 7.43-7.46 (m, 1H); ¹³CNMR: (125 MHz, CDCl₃) δ: 162.40, 153.51, 148.91, 135.79, 132.39, 131.81, 130.93, 128.10, 126.59, 125.87, 124.61, 123.94, 121.58.

17. 4-(1,3-benzothiazol-2-yl)-N, N-dimethylaniline (Table 6, Entry 17, 7q):White solid $161-163^{\circ}C$ ($160-162^{\circ}C$)¹⁴



18. 2-[(E)-2-phenylethenyl]-1,3-benzothiazole (Table 6, Entry 18, 7r):White solid M. P. 107-110°C (110-112°C)¹⁴



19. 2-(furan-2-yl)-1,3-benzothiazole (Table 6, Entry 19, 7s): White solid M. P. 103-104°C (101-102°C)¹⁰



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¹HNMR: (500 MHz, CDCl₃): δ 8.04 (d, 1H, *J*=8.04 Hz), 7.88 (d, 1H, *J*=7.88 Hz), 7.60 (d, 1H, *J*=7.59 Hz), 7.47-7.50 (m, 1H), 7.36-7.39 (m, 1H), 7.18 (d, 1H, *J*=7.18 Hz), 6.59-6.60 (m, 1H); ¹³CNMR: (125 MHz, CDCl₃)δ:157.56, 153.74, 148.73, 144.70, 134.26, 126.48, 125.19, 123.11, 121.57, 112.53, 111.43.

20. 2-(pyridin-2-yl)-1,3-benzothiazole (Table 6, Entry 20, 7t):Brown solid 132-134°C (130-132°C)¹⁶



¹HNMR: (500 MHz, CDCl₃): δ 8.67-8.68 (m, 1H), 8.36 (d, 1H, *J*=8.35 Hz), 8.08 (d, 1H, *J*=8.08 Hz), 7.94 (d, 1H, *J*=7.94 Hz), 7.80-7.84 (m, 1H), 7.47-7.51 (m, 1H), 7.34-7.42 (m, 2H); ¹³CNMR: (125 MHz, CDCl₃) δ : 169.35, 154.25, 151.36, 149.63, 136.99, 136.09, 126.26, 125.63, 125.25, 123.55, 122.00, 120.73. **21.** 2-(1*H*-indol-2-yl)-1,3-benzothiazole (Table 6, Entry 21, 7u):Brown solid M. P. 146-148°C (144-147°C)¹⁶



¹HNMR:(**500** MHz, CDCl₃):δ 8.82 (s, 1H), 8.44 (d, 1H, *J*=8.43 Hz), 8.03 (d, 1H, *J*=8.03 Hz), 7.93 (d, 1H, *J*=7.92 Hz), 7.88 (d, 1H, *J*=7.87 Hz), 7.46 δ (t, 1H, *J*=7.46 Hz), 7.43 (t, 1H, *J*=7.41 Hz) 7.35-7.28 (m, 3H); ¹³CNMR: (**125** MHz, CDCl₃) δ: 163.00, 153.730, 136.46, 133.84, 126.34, 126.07, 124.92, 124.23, 123.44, 122.11, 121.83, 121.30, 121.05, 112.46, 111.67.

22: 3-(1,3-benzothiazol-2-yl)-6-bromo-4*H*-1-benzopyran-4-one. (Table 6, Entry 22, 7v): Yellow solid M. P. 254-256 °C.



'HNMR: (500 MHz, CDCl₃): δ 9.28 (s, 1H), 8.50 (d, 1H, *J*=8.49 Hz), 7.98-8.04 (m, 2H), 7.83 (q, 1H), 7.46-7.53 (m, 2H), 7.41 (t, 1H, *J*=7.40 Hz); ¹³CNMR:(125 MHz, CDCl₃) δ : 173.58, 158.03, 156.55, 154.69, 151.63, 137.33, 136.08, 128.98, 126.32, 125.17, 124.97, 122.57, 121.68, 120.34, 119.75, 118.52. HRMS: [MF: C₁₆ H₉ O₂ NS Br(M+H)]: 359.95 (Calculated: 359.21).




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Graph: 1 (1a) DBU-iodine complex UV-visible spectrum.

S38



Accepted manuscrity

S41

Graph:2 (1b) Morpholine-iodine complex UV-visible spectrum.



S43

Graph: 3. (1c) Urotropine-iodine complex UV-visible spectrum.



S45

Graph: 4 (1d) Piperazine-iodine complex UV-visible spectrum.



S47 A HI CHINGS CHIN

Graph: 5 (1e) N-methyl-Piperazine-iodine complex UV-visible spectrum.

S48



Graph: 1 (1a) DBU-iodine complex IR spectrum.

S50



Accepted manuscrity

Graph:2 (1b) Morpholine-iodine complex IR spectrum.

S53


A HIGHLING CALL

Graph: 3. (1c) Urotropine-iodine complex IR spectrum.



Accepted manuscrity

Graph: 4 (1d) Piperazine-iodine complex IR spectrum.



S61

Graph: 5 (1e) N-methyl-Piperazine-iodine complex IR spectrum.



A HIGH MASSING GAWADE and KULKARNI.





S64



GAWADE and KULKARNI.



S66



Figure: 5 (1e) SEM of N-Methyl-piperazine-iodine complex indicate Morphology.

GAWADE and KULKARNI.



Field Emission Scanning Electron Microscopy Energy Dispersive X-ray Spectroscopy (*FESEM* - EDS): Figure: 1 (1a) DBU-iodine

complex.



S70 GAWADE and KULKARNI. Graph: 1 (1a) DBU-iodine complex. cps/eV 5-4-3-N CI 2-1-0 3 2 5 1 6 4 keV

Figure: 2 (1b) Morpholine-iodine complex.



GAWADE and KULKARNI.







S73

Graph: 3 (1c) Urotropine Iodine Complex.







GAWADE and KULKARNI.





Figure: 5 (1e) N-methyl-piperazine-iodine complex.

S76



S77



Graph: 1(1a) TGA-DTA of DBU-iodine complex.



Graph: 2(1b) TGA-DTA of Morpholine-iodine complex.



Graph: 3(1c) TGA-DTA of Urotropine-iodine complex.



Graph: 4(1d) TGA-DTA of Piperazine-iodine complex.







S83

Fig: ¹H-NMR DBUH-I₃ complex (Table 1, Entry 1, 1a)



S85

Fig: ¹³C-NMR DBUH-I₃ complex (Table 1, Entry 1, 1a)



S87

Fig: ¹H-NMR Morpholine-Iodine complex (Table 1, Entry 2, 1b)



S89

Fig: ¹³C-NMR Morpholine-Iodine complex (Table 1, Entry 2, 1b)


SUPPLEMENTARY MATERIAL

S91

Fig: ¹H-NMR Urotropine-Iodine complex (Table 1, Entry 3, 1c)



SUPPLEMENTARY MATERIAL

S93

Fig: ¹³C-NMR Urotropine-Iodine complex (Table 1, Entry 3, 1c)



SUPPLEMENTARY MATERIAL

S95

Fig: ¹H-NMRPiperazine-Iodine complex (Table 1, Entry 4, 1d)



SUPPLEMENTARY MATERIAL

S97

Fig: ¹³C-NMRPiperazine-Iodine complex (Table 1, Entry 4, 1d)



Fig: ¹³C-NM R N-Met hyl-Pipe razi

ne-



Iodine complex (Table 1, Entry 5, 1e)



Fig. ¹H-NMR of 2-(4-chlorophenyl)-*1H*-benzimidazole (Table 5, Entry 1, 4a)



Fig. ¹³C-NMR of 2-(4-chlorophenyl)-1H-benzimidazole (Table 5, Entry 1, 4a)



Fig:¹H-NMR2-(3-chlorophenyl)-1*H*-benzimidazole (Table 5, Entry 2, 4b)



Fig: ¹³C-NMR 2-(3-chlorophenyl)-1*H*-benzimidazole (Table 5, Entry 2, 4b)



Fig: ¹H-NMR 2-(2-chlorophenyl)-*1H*-benzimidazole (Table 5, Entry 3, 4c)



SUPPLEMENTARY MATERIAL

Fig: ¹³C-NMR 2-(2-chlorophenyl)-*1H*-benzimidazole (Table 5, Entry 3, 4c)



Fig: ¹H-NMR 2-(4-Methoxyphenyl)-*1H*-benzimidazole (Table 5, Entry 5, 4e)



SUPPLEMENTARY MATERIAL

SUPPLEMENTARY

Fig: ¹³C-NMR 2-(4-Methoxyphenyl)-*1H*-benzimidazole (Table 5, Entry 5, 4e)



Fig: ¹H-NMR 2-(2H-1,3-benzodioxol-5-yl)-*1H*-benzimidazole (Table 5, Entry 8, 4h)



SUPPLEMENTARY MATERIAL

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-1

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125.7763643 MHz 130 8.25 User 100.00000000 W 500.130005 MHz 1H Maltz16 8.00 User 22.0000000 W 0.2922000 W 0.14698000 W

Fig: ¹³C-NMR 2-(2H-1,3-benzodioxol-5-yl)-1H-benzimidazole (Table 5, Entry 8, 4h)



Fig: ¹H-NMR2-phenyl-1*H*-benzimidazole (Table 5, Entry 9, 4i)



Fig: ¹³C-NMR2-phenyl-*1H*-benzimidazole (Table 5, Entry 9, 4i)


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Fig: ¹H-NMR 4-(*1H*-benzimidazole-2-yl) phenol (Table 5, Entry 11, 4k)



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Fig: ¹³C-NMR 4-(1*H*-benzimidazole-2-yl) phenol (Table 5, Entry 11, 4k)



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Fig: ¹H-NMR 2-(*1H*-benzimidazole-2yl) phenol (Table 5, Entry 12, 4l)



Fig: ¹³C-NMR 2-(*1H*-benzimidazole-2yl) phenol (Table 5, Entry 12, 4l)



Fig: ¹H-NMR 2-(2-nitrophenyl)-*1H*-benzimidazole (Table 5, Entry 15, 40)



Fig: ¹³C-NMR 2-(2-nitrophenyl)-*1H*-benzimidazole (Table 5, Entry 15, 40)



Fig: ¹H-NMR 4-(*1H*-benzimidazole-2-yl)-N, N-dimethylaniline (Table 5, Entry 16,4p)



Fig: ¹³C-NMR 4-(*1H*-benzimidazole-2-yl)-N, N-dimethylaniline (Table 5, Entry 16,4p)



Fig: ¹H-NMR 3-(*1H*-benzimidazole-2-yl)2-chloroquinoline (Table 5, Entry 19, 4s)



Fig: ¹³C-NMR 3-(*1H*-benzimidazole-2-yl)2-chloroquinoline (Table 5, Entry 19, 4s)



Fig: ¹H-NMR 3-(*1H*-benzimidazol-2-yl)-2-chloro-6-methylquinoline (Table 5, Entry 20, 4t)



Fig: ¹³C-NMR 3-(*1H*-benzimidazol-2-yl)-2-chloro-6-methylquinoline (Table 5, Entry 20, 4t)



Fig: ¹H-NMR 2-(1*H*-indol-3-yl)-*1H*-benzimidazole (Table 5, Entry 21, 4u)



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Fig: ¹³C-NMR 2-(1*H*-indol-3-yl)-*1H*-benzimidazole (Table 5, Entry 21, 4u)



Fig: ¹H-NMR 3-(*1H*-benzimidazol-2-yl)-6-bromo-4*H*-1-benzopyran-4-one (Table 5, Entry 22, 4v)



Fig: ¹³C-NMR 3-(*1H*-benzimidazol-2-yl)-6-bromo-4*H*-1-benzopyran-4-one (Table 5, Entry 22, 4v)

In (Table 5, Entry 22, 4)





Fig: ¹H-NMR 2-(4-bromophenyl)-1,3-benzothiazole (Table 6, Entry 4, 7d)


Fig: ¹³C-NMR 2-(4-bromophenyl)-1,3-benzothiazole (Table 6, Entry 4, 7d)



Fig: ¹H-NMR 2-(4-methoxyphenyl)-1,3-benzothiazole (Table 6, Entry 5, 7e)



Fig: ¹³C-NMR 2-(4-methoxyphenyl)-1,3-benzothiazole (Table 6, Entry 5, 7e)



Fig: ¹H-NMR 2-(2*H*-1,3-benzodioxol-5-yl)-1,3-benzothiazole (Table 6, Entry 8, 7h)



Fig: ¹³C-NMR 2-(2*H*-1,3-benzodioxol-5-yl)-1,3-benzothiazole (Table 6, Entry 8, 7h)



Fig: ¹H-NMR 4-(1,3-benzothiazol-2-yl) phenol (Table 6, Entry 11, 7k)



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Fig: ¹³C-NMR 4-(1,3-benzothiazol-2-yl) phenol (Table 6, Entry 11, 7k)



Fig: ¹H-NMR 2-(1,3-benzothiazol-2-yl) phenol (Table 6, Entry 12, 7l)



Fig: ¹H-NMR 2-(1,3-benzothiazol-2-yl) phenol (Table 6, Entry 12, 7l)



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Fig: ¹H-NMR 4-(1,3-benzothiazol-2-yl)-2-methoxyphenol (Table 6, Entry 13, 7m)



Fig: ¹³C-NMR 4-(1,3-benzothiazol-2-yl)-2-methoxyphenol (Table 6, Entry 13, 7m)



Fig: ¹H-NMR 2-(4-nitrophenyl)-1,3-benzothiazole (Table 6, Entry14, 7n)



Fig: ¹³C-NMR 2-(4-nitrophenyl)-1,3-benzothiazole (Table 6, Entry14, 7n)



Fig: ¹H-NMR 2-(2-nitrophenyl)-1,3-benzothiazole (Table 6, Entry 16, 7p)



Fig: ¹³C-NMR 2-(2-nitrophenyl)-1,3-benzothiazole (Table 6, Entry 16, 7p)



Fig: ¹H-NMR 4-(1,3-benzothiazol-2-yl)-*N*, *N*-dimethylaniline (Table 6, Entry 17, 7q)



Fig: ¹³C-NMR 4-(1,3-benzothiazol-2-yl)-*N*, *N*-dimethylaniline (Table 6, Entry 17, 7q):



Fig: ¹H-NMR2-(furan-2-yl)-1,3-benzothiazole (Table 6, Entry 19, 7s)


Fig: ¹H-NMR2-(furan-2-yl)-1,3-benzothiazole (Table 6, Entry 19, 7s)



Fig: ¹H-NMR 2-(pyridin-2-yl)-1,3-benzothiazole (Table 6, Entry 20, 7t)



Fig: ¹³C-NMR 2-(pyridin-2-yl)-1,3-benzothiazole (Table 6, Entry 20, 7t)



Fig: ¹H-NMR 2-(1*H*-indol-3-yl)-1,3-benzothiazole (Table 6, Entry 21, 7u)



Fig: ¹³C-NMR 2-(1*H*-indol-3-yl)-1,3-benzothiazole (Table 6, Entry 21, 7u)



Fig: ¹H-NMR 3-(1,3-benzothiazol-2-yl)-6-bromo-4*H*-1-benzopyran-4-one. (Table 6, Entry 22, 7v)



Fig: ¹³C-NMR 3-(1,3-benzothiazol-2-yl)-6-bromo-4*H*-1-benzopyran-4-one. (Table 6, Entry 22, 7v)

SUPPLEMENTARY MATERIAL

Fig: 3-(1,3-benzothiazol-2-yl)-6-bromo-4H-1-benzopyran-4-one. (Table 6, Entry 22, 7v)

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